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## Galvanic cell experiment lab report

Academia.edu uses cookies to personalize content, adapt ads, and improve user experience. Using our website, you agree to our collection of information using cookies. To find out more, check out our privacy policy. x SULTAN IDRIS LABORATORY REPORT SKU3023: CHEMISTRY II Semester II Session 2016/2017 ID NO AND NO 1. Abdul Rahim bin Md Saad (D20162075593) 2. Mohd Hafiz Ayman bin Abdullah (D20162075602) Teacher Professor Madya Dr. Ismail zainol EXPERIMENT NUMBER 4 Name Electrochemistry DATA and DAY 23 Rd October 2017, Monday / 09:05, Monday / 09: CHECK THE LIST (Please tick off) Name Goal (s) Acquaintance Methods (on the chart) Observations (in table) Data and calculation Discussion Questions and answers (if necessary) Conclusion (s) Links (at least 2) EXPERIMENT 4: ELECTROCHEMIA 1.0 Goal 1.1. Explain the concept of the flow of electrons, anions and cations. 1.2. Identify the relative potential of reducing redox reactions. 1.3. To describe the effect of concentration on cellular potential. 2.0 Acquaintance Electrochemical reaction, any process caused or accompanied by the passage of an electric current and involving in most cases the transmission of electrons between two substances — one is solid and the other is liquid. (Bocris and Despic, 2011) The oxidation reaction occurred in the anode and the reaction decrease occurred in the cathode Since the transmission of electrons occurs spontaneously, one set of galvanic cells can be developed by separating the vapor of redox. The transmission of electrons is no longer in the inter-phases, but through an external circuit. The two halves of the redox reaction are divided into two different containers (such as a glass) and are called semi-cellular. The general view of the semi-flight consists of a piece of metal (electrode) immersed in the solution of its ions. The electrons are wired. The scheme between the two solutions ends with a salt bridge. The salt bridge allows electrical contact between two semi-cells, the inner chain. It maintains electrical neutrality in each half of the cell as ions stick in and out of the salt bridge. The pored pot also serves the same function as a salt bridge. 3.0 ACTIVITIES Section A: Galvanic cell chemicals Machine Copper Metal Zinc Metal Magnesium Metal 0.1M Copper Sulfate Solution 0.1M zinc salt solution 0.1M magnesium saline solution 0.1M iron salt solution 50 ml glass wire sand paper Voltmeter filter paper filter Procedures Link to jotter 1.0 Section B: Cell Potential Chemicals Machine zinc metal Copper metal 0.1M CuSO 4 0.01M CuSO solution 4 0.001M CuSO solution 4 6 M NH3 Solution 0.2M Na2S Solution 50ml glass of wire Sand paper Voltmeter filter paper Procedures Link to Jotter 2.0

